



(29)

(ai)

Steel

(aii)

aluminium can be used in oxidising conditions because aluminium can undergo the physical reaction that happens from oxidising conditions and doesn't change quickly.

(bi)

Iron

(bii)

Sacrificial anodes are added to metal-hulled ships as they have slow corrosion levels, as they take a long time to rust which is better when they are out to sea.

(c)

The effects from adding other elements to iron on the properties and uses of steel as that



you would have a bigger problem with rust as steel tends to rust very quickly, also by mixing the elements you might lose the iron's strength and reliability as steel tends to break a lot quickly.

(di)

corrosion is a form of rust.

(dii)

One procedure a school laboratory can use to monitor corrosion levels is to have a piece of each ~~metal~~ metal, set them in their own beaker add water and all that's left is observation.

(diii)

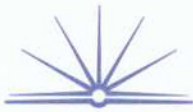
ways in which accuracy and reliability could be improved is to try and speed up the process as you could then



witness the changing taking place instead of hoping it wont happen whilst your at home or in another class.

(e)

steel that is found from recovered shipwrecks cant just be brought up to the surface, allowed to dry by the oxygen as then the salt would dry upon the metal leaving it, ~~leaving~~ in a bad misformed shape, what you need to do is bring it up to the surface but place it in some  $H_2O$  and allow the metal to be cleaned, once cleaned you can then take it out and allow to dry normally as the salt would be gone, but then you need to think of rust cause by this time the rust might have already sunk in and making its way through the whole metal, hoping that when you fish the



Shipwreck from the water that it has some  
sort of protection on the steel.